



Section 5.2

Salts

Study Notes

By the end of section 5.2 you should be able to understand the following:

- Salts are compounds that include ionic compounds that form when acids and bases react.
- When acids and bases react, the process is called neutralization.
- Metals reacting with acids, and oxides or carbonates reacting with acids, can also form salts.
- Metal oxides react with water to form bases, and non-metal oxides react with water to form acids.

NOTES

Where is normal table salt normally obtained from?

- 1.
- 2.
- 3.

In chemistry, what is the definition of a salt?

- 1.

Give two examples of acid-base neutralizations that produce a salt, and label the salt by circling it in your answer.

- 1.
- 2.

Give an example of a metal oxide reacting with water to form a base, Circle the base in your answer.

- 1.

Give an example of a non-metal oxide reacting with water to form an acid. Circle the acid in your answer.

- 1.

Do the Reading Check on page 238

NOTES

Where are the most reactive metals on the periodic table found?

1.

Give an example of a highly reactive metal reacting with an acid to form a salt. Circle the salt in your answer.

1.

Give an example of a less reactive metal reacting with an acid to form a salt. Circle the salt in your answer.

1.

What are the two salts formed when sulphuric acid and nitric acid, found in acid precipitation, react with carbonates?

1.

2.