Section 5.1
Acids and Bases
Study Notes

By the end of section 5.1 you should be able to understand the following:

- ❒ Many common compounds can be classified as acids or bases.
- ❒ Indicators are compounds that reveal the level of acidity of other compounds.
- ❒ The pH scale is a numerical method that indicates how acidic or basic a compound is.
- ❒ Acids generally have a chemical formula starting with an H-, whereas bases generally end with an -OH

NOTES

What are some common examples of acids and bases?	1. 2.
What safety issues must be remembered when dealing with acids and bases?	1. 2. 3.

NOTES

List some common acids, neutral substances and bases	1. 2. 3.
How much greater is the acidity of a substance with pH = 4 than a neutral substance?	1.
What are six different methods for measuring pH?	1. 2. 3. 4. 5. 6.
What are three indicators that could be used to determine the pH of an unknown acid?	1. 2. 3.

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NOTES

List some common qualities of acids

1.

2.

3.

What are the rules for naming acids ending in
–***ide*** *(hydrogen chloride)*
–***ate*** *(hydrogen carbonate)*
and
-***ite*** *(hydrogen sulfite)*?

1.

2.

3.

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What is an easy way to identify a base from its chemical formula?

1.

What is the relationship between H^+ and OH^- ions in acids and bases?

1.

2.

3.

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NOTES

List some common properties of acids and bases

Properties	Acids	Bases
Taste		
Touch		
Indicator tests		
Rxn w/ metals		
Conduct electricity		
pH		
Ions produced		